

Honors Acid Bases ... Bronsted Lowery Conjugate Acids and Bases

Complete the following:

What is an Acid in the Bronsted-Lowery model:

Donates a proton (H^+)

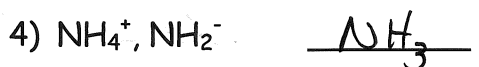
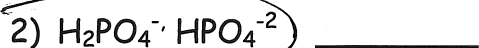
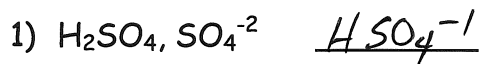
What is an Base in the Bronsted-Lowery model:

Accepts a proton (H^+)

When acid is dissolved in water, what ion does the water form:

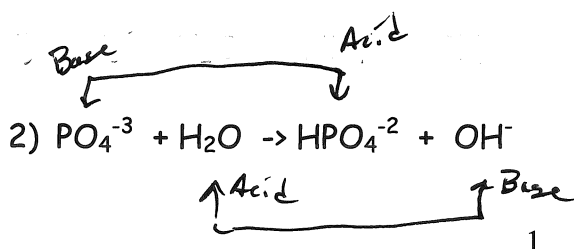
H_3O^+ Hydronium

Which of the following represent conjugate acid base pairs (circle them):



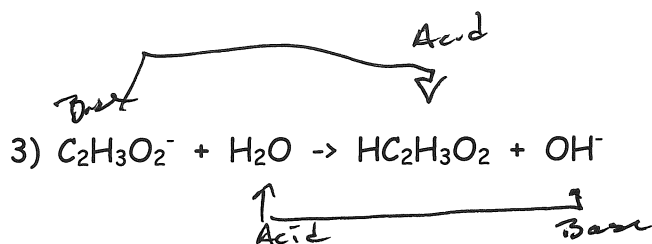
Correct the above incorrect pairs by changing the 2nd formula to be the correct conjugate acid or base.

In each of the following chemical equations, identify the conjugate acid base pairs.



Base solutions

Base solution



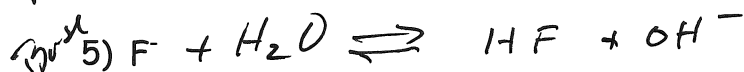
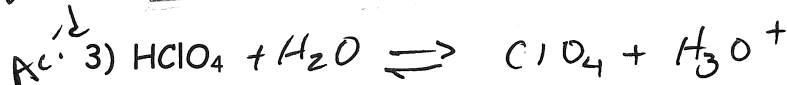
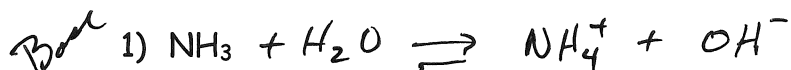
Write the conjugate acid for each of the following:



Write the conjugate base for each of the following:



Write the chemical equation showing how each of the following behaves in water.



O^{2-}
 $\uparrow$
Corrections!